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**An experimental and theoretical investigation on the optical and photocatalytic properties of ZnS nanoparticles**

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### Abstract

From the viewpoints of materials chemistry and physical chemistry, crystal structure directly determines the electronic structure and furthermore their optical and photocatalytic properties. Zinc sulfide (ZnS) nanoparticles (NPs) with tunable photoluminescence (PL) emission and high photocatalytic activity have been obtained by means of a microwave-assisted solvothermal (MAS) method using different precursors (i.e., zinc nitrate (ZN), zinc chloride (ZC), or zinc acetate (ZA)). The morphologies, optical properties, and electronic structures of the as-synthesized ZnS NPs were characterized by X-ray photoelectron spectroscopy (XPS), transmission electron microscopy (TEM), energy dispersive X-ray spectroscopy (EDX), Brunauer–Emmett–Teller (BET) isotherms for N<sub>2</sub> adsorption/desorption processes, diffuse reflectance spectroscopy (DRS), PL measurements and theoretical calculations. Density functional theory calculations were used to determine the geometries and electronic properties of bulk wurtzite (WZ) ZnS NPs and their (0001), (10 $\bar{1}$ 0), (11 $\bar{2}$ 0), (10 $\bar{1}$ 1), and (10 $\bar{1}$ 2) surfaces. The dependence of the PL emission behavior of ZnS NPs on the precursor was elucidated by examining the energy band structure and density of states. The method for degradation of Rhodamine B (RhB) was used as a probe reaction to investigate the photocatalytic activity of the as-synthesized ZnS NPs under UV light irradiation. The PL behavior as well as photocatalytic activities of ZnS NPs were attributed to specific features of the structural and electronic structures. Increased photocatalytic degradation was observed for samples synthesized using different precursors in the following order: ZA<ZC<ZN. These results indicated that samples synthesized with ZN present a greater percentage of exposed (0001) surface than those synthesized with the ZC and ZA. Furthermore, the possible photodegradation mechanism of the as-prepared ZnS NPs were also briefly discussed.

**Keywords:** ZnS; Nanoparticles; Microwave-assisted solvothermal method; DFT; Photocatalysis; Photoluminescence

## Introduction

The interest in ZnS NPs has increase mainly because of their unique properties and enormous potential in emerging technologies ranging from optoelectronic luminescent devices to photocatalysis [1-8]. It is known that bulk ZnS is an important binary semiconductor with a direct, wide bandgap (~3.6 eV), high refraction index, and high transmittance in the visible range [9,10]. A variety of synthetic strategies has been used to obtain ZnS crystals with a superior control of their physical characteristics (e.g., sizes, well-defined morphologies, and compositions), including hot injection [11,12], solid state reaction [13], thermal decomposition [14], co-precipitation [15,16], hydro/solvothermal method [13, 17-21] and son on. **On the other hand, the doping with impurities in ZnS NPs can display an increased efficiency in heterogeneous catalysis process, as well as in its optical properties [22-24].** However, despite recent advances in the chemical methods of synthesis, it is mandatory to conduct improvements with respect the existing synthetic methods in order to develop more efficient and environmentally friendly methods that allow the ability to synthesize these new complex materials with various functionalities and desired properties, as basis for the future development of new technologies at the nanoscale. In this context, more recently, our research group has employed the microwave-assisted solvothermal (MAS) method as a promising strategy to obtain ZnS crystals at shortened reaction times, low-temperatures and, ultimately, cleanness reactions compared to conventional routes [4,9,21,25]. **Hence, there is one large interest in the development of the novel NPs based on ZnS with its tailored surfaces exposed to a variety of technological applications at the nanoscale [4, 22-29].**

Intrinsic and visible emissions in ZnS NPs have been well-reported [30-32], and their PL properties have been explored by many researchers [9,25,33-36], while the

photocatalytic properties of ZnS NPs have been analyzed by many authors for a variety of reactions [27-29,37-42], however, the photocatalytic and PL behavior at nanoscale sizes remain far from well-understood.

First-principles calculations, mainly within the DFT framework, have proved invaluable in providing guidance and an atomistic understanding of the physical and chemical properties of emerging complex functional nanomaterials. Within this general framework, we are engaged in a joint experimental and theoretical project that is concerned with the synthesis, characterization, and technological applications of a series of Pb and Ag-based metal oxides, such as  $\text{PbMoO}_4$  [43],  $\alpha\text{-Ag}_2\text{WO}_4$  [44],  $\beta\text{-Ag}_2\text{WO}_4$  [45],  $\beta\text{-Ag}_2\text{MoO}_4$  [46],  $\beta\text{-AgVO}_3$  [47],  $\text{Ag}_2\text{CrO}_4$  [48] or Ag-containing salts such as  $\text{Ag}_3\text{PO}_4$  [49,50].

Electronic and structural properties of the ZnS NPs as well as their surfaces can be revealed straightforwardly by means of theoretical calculations and have been studied extensively in the literature [51-57]. For instance, Barnard *et al.* [52] reported the theoretical equilibrium morphologies of the ZnS NPs in the wurtzite (WZ) phase as a function of size, as determined by DFT calculations. Qin *et al.* [53] reported the photocatalytic efficiency of the zinc blend and WZ phases using first-principles DFT calculations for both the bulk samples and at the polar surfaces. Furthermore, the effect of water on crystallinity enhancement as well as on the structure transformation of ZnS NPs at room temperature has been reported in other theoretical studies [55-57]. More recently, we reported the growth mechanism of ZnS NPs [4], and a DFT study of the structural and electronic properties of ZnS polymorphs and their pressure-induced phase transitions [51].

In this paper, a systematic investigation, based on the joint use of experimental data and theoretical calculations, into the morphologies, electronic structures, optical

properties and photocatalytic activities of ZnS NPs has been presented. Three different precursors (i.e., ZN, ZC, or ZA) have been employed along the synthesis performed by means of the MAS method. These ZnS NPs were analyzed by XPS, TEM, EDX, BET isotherms for N<sub>2</sub> adsorption/desorption processes, DRS, and PL measurements. Photocatalytic activities were evaluated via degradation of RhB under UV irradiation. DFT calculations allows us to carry out a detailed analyses of the electronic structures of both bulk WZ ZnS NPs and the corresponding (0001), (10 $\bar{1}$ 0), (11 $\bar{2}$ 0), (10 $\bar{1}$ 1), and (10 $\bar{1}$ 2) surfaces of the WZ ZnS and the origins of the intense PL bands and photocatalytic activities of ZnS were elucidated.

The rest of the paper is divided into three sections. Next, we present the experimental and theoretical methods and, our results presented are discussed in the section following the next. Main conclusions are summarized in the last section.

### **Experimental and theoretical methods**

All reagents were analytical grade and used without further purification. ZnS NPs were prepared using the MAS method at 140 °C for 10 min using three different zinc salt precursors, i.e., ZN, ZC, or ZA, in accordance with a protocol developed in a previous study [9]. Briefly, the synthesis of ZnS is described as follows: initially, 7.34 mmol of precursors (i.e., ZN, ZA or ZC) and 15.44 mmol of tetrabutylammonium hydroxide were dissolved in 25 mL of ethylene glycol (EG) and heated to 80 °C (solution 1). Afterward, 7.34 mmol of thiourea were separately dissolved in another 25 mL of EG (solution 2. After some minutes, under vigorous magnetic stirring, the solution 2 was then quickly mixed in the solution 1. Then, the mixture was transferred into a Teflon autoclave which was sealed and placed inside a domestic microwave-

solvothermal system (2.45 GHz, maximum power of 800 W) at 140 °C for 10 min. The resulting white solid was washed with deionized water and ethanol to remove byproducts possibly remaining in the final product and the precipitates were finally collected and dried at 70 °C overnight. Based on this protocol, the XRD results confirm that all diffraction peaks can be indexed to the hexagonal structure of ZnS and the samples as-prepared have a high purity and crystallinity [9].

The obtained powders were structurally characterized using XPS spectra were collected using a commercial spectrometer (UNISPECS-UHV) using the Mg K $\alpha$  line ( $h\nu = 1253.6$  eV). The composition of the surface region was determined from the ratio of the relative peak areas corrected by the sensitivity factors (Scofield) of the corresponding elements. Spectra were fitted without placing constraints using multiple Voigt profiles, and the BEs were corrected using the hydrocarbon component of adventitious carbon fixed at 285.0 eV. The morphologies were investigated using TEM images, which were recorded on a JEOL 2100 transmission electron microscope operated at 200 kV. During sample preparation for the TEM technique, the suspensions were deposited onto copper grids via fast immersion. BET isotherms for N<sub>2</sub> adsorption/desorption processes were obtained by using a Micromeritics ASAP 2010 surface area and porosity analyzer.

The DRS spectra in the ultraviolet-visible region were recorded at room temperature between 200 and 800 nm using a Varian model Cary 5G in the diffuse reflectance mode. PL spectra were collected using a Thermal Jarrel-Ash Monospec monochromator and Hamamatsu R446 photomultiplier. The 350.7 nm (2.57 eV) exciting wavelength of a krypton ion laser (Coherent Innova) was used with the output of the laser maintained at 200 mW. All measurements were performed at room temperature.

Photocatalytic activity of ZnS NPs synthesized with different zinc salts precursors

was tested in the degradation of RhB under UV. RhB is used as a model compound for oxidation reactions and has a strong absorption in the visible region ( $\lambda_{\text{max}} = 554 \text{ nm}$ ). RhB is highly soluble in water and has properties similar to those of textile dyes, which are difficult to degrade. RhB dye solution (80 mL,  $10 \text{ mgL}^{-1}$ ) at pH 6.0 was mixed with 60 mg of the catalyst and irradiated with an ultraviolet light placed in a box, which was set at a distance of 20 cm from the beaker containing the suspension. These systems remained under constant stirring in the dark for a period of 15 minutes to achieve adsorption-desorption equilibrium on the surface of the ZnS NPs. After this time, the first aliquot was collected and considered as time zero. Photocatalytic tests were performed at  $20 \text{ }^\circ\text{C}$  in thermostatic bath with vigorous stirring. The UV radiation used in this experiment was obtained from a medium-pressure mercury vapor lamp (Osram, HQL 400) using the output without bulb protection. Reactions were monitored by UV-vis spectroscopy (JASCO V-660) at 554 nm using a commercial quartz cuvette.

To complement the experimental data, first-principles calculations were performed within the periodic DFT framework using the VASP program [58,59]. The Kohn-Sham equations were solved using the Perdew, Burke, and Ernzerh exchange-correlation functional [60], and the electron-ion interaction was described by the projector-augmented-wave pseudopotentials [58,61]. The plane-wave expansion was truncated at a cut-off energy of 520 eV and is required for different convergence precisions, and the Brillouin zones were sampled through Monkhorst-Pack special k-points [62] grids ( $6 \times 6 \times 1$ ) that ensure geometrical and energetic convergence for the ZnS structures considered in this work. A vacuum spacer of  $15 \text{ \AA}$  was introduced in the z-direction to prevent the surfaces from interacting with each other. The polar (0001) and semi-polar ( $10\bar{1}1$ ) surfaces were modeled with the top Zn layer and bottom S layer (each S atom with one dangling bond) using a  $2 \times 2 \times 1$  conventional supercell with 20 atomic layers.

The  $(10\bar{1}0)$ ,  $(11\bar{2}0)$  and  $(10\bar{1}2)$  facets were modeled by slabs containing twelve ZnS layers.

## Results and Discussion

High-resolution XPS spectra of the Zn 2p and S 2p orbitals for the WZ ZnS NPs are presented in Figure 1. The high-resolution Zn 2p spectrum shows BEs around 1020, 1022, and 1023 eV, and these values can be associated, respectively, with the presence of Zn–C, Zn–S, and Zn–O bonds at the surface, which is in accordance with other studies [9]. Moreover, it is important to remark that XPS results reveals a small amount of Zn-C and Zn-O bonds with respect Zn-S bonds present in the XPS spectrum of all samples, which can be attributed to an insignificant contamination by such particles, i.e., hence the ZnS NPs prepared by the MAS method presents a well-cleaned surface and confirmed a high purity for such particles. The fitted S 2p spectra peaks (see Figure 1) at approximately 161 and 163 eV were assigned to ZnS  $2p_{3/2}$  and ZnS  $2p_{1/2}$  orbitals, respectively [9,63,64]. No signs of a secondary phase were observed using XPS spectra (this result is confirmed by TEM analysis, see supplementary material) for the ZnS NPs synthesized using different precursors. These results are consistent with BE values reported in the literature [63-65], and confirm the purity of ZnS NPs prepared using the MAS method.

On the other hand, the variations observed in the BEs are associated with the concentrations of sulfur defects on the surfaces of the ZnS NPs. Three different charge states of sulfur vacancies are evident:  $V_S^x$ ,  $V_S^\bullet$ , and  $V_S^{\bullet\bullet}$ ; these play a key role in understanding the physical and chemical properties of these NPs [9,63-67]. In addition, surface composition analysis via XPS measurements revealed ratios of Zn/S atoms of



1:1, 1:0.9, and 1:0.8 for the ZnS NPs synthesized from ZN, ZC, or ZA, respectively.

### Figure 1.

The adsorption and desorption isotherm curves of the ZnS NPs generated from the three precursors are shown in the Supplementary Information. The specific surface areas were 233, 283, and 295  $\text{m}^2\text{g}^{-1}$  for the ZnS NPs produced from ZN, ZC, or ZA precursors, respectively. The results showed Type I isotherms, which indicates the predominance of mesopores in the three synthesized samples. These results are in good agreement with literature values [68]. Owing to the high surface areas of the ZnS NPs prepared using the MAS method, it is likely that these materials formed from small particles or clusters rather than single-crystalline ZnS, which is in good agreement with our TEM results (see Supporting Information). Furthermore, the aggregation state for such NPs reveal a strong polarization in the ZnS surfaces and can in principle provide a rougher surface.

The optical properties were analyzed using DRS spectroscopy and PL measurements (see Figures 2 (a) and (b)). In particular, the DRS spectra in Figure 2 (a) shows variation in the optical direct band gaps of the ZnS NPs from 3.45 to 3.54 eV. For details on the calculations of the optical band gap see reference [9]. The band gap values estimated for the samples prepared using different precursors along the MAS synthesis are much lower than that expected for the band gap values of 3.77 eV at room temperature for bulk WZ ZnS reported by Fang *et al.* [2]. Along the text, the term “defect” it is commonly used here in its broadest sense to encompass a disturbance in the ideal crystal framework, e.g., this definition may include dislocations, vacancies, presence of the impurities, even the surface of the material, and so on [69-71].

The variation in the DRS spectra is controlled by the degree of structural and electronic disorder on the crystalline phase of the ZnS NPs. This fact reveals that this disordered structure promotes local polarization and a charge gradient in the structure, which is consistent with the XPS results. Therefore, we believe that the structural changes provoked by the use of the different types of precursors can cause a variation in the optical band gap of the ZnS NPs, and hence is responsible for modulation of its physical and chemical properties, i.e., leading to different PL and photocatalytic behaviors of such NPs.

### Figure 2.

Figure 2 (b) shows the PL spectra at room temperature of ZnS NPs synthesized using the MAS process with different zinc salts. The PL spectra feature broad bands covering the visible electromagnetic spectrum from 400 to 800 nm, with maximum emissions at 478, 505, and 555 nm for the ZnS NPs synthesized using ZN, ZC, or ZA precursors, respectively, when excited by a 350.7 nm laser line. An analysis of the results points out the presence of red-shift of 0.36 eV, due to the structural changes in the ZnS NPs. This behaviour, in principle, offers an excellent platform for the design and control of ZnS NPs that emit different colours via band-gap engineering.

An analysis of the PL emissions renders certain localized states that must exist in the forbidden band gap [4,9]. It has been reported that the PL emission of ZnS NPs in the visible region is due to sulfur vacancies ( $V_S^x$ ,  $V_S^+$ , and  $V_S^{2+}$ ) that generate shallow and deep levels within the band gap [9,24,25]. Thus, a blue-shift in PL emission correlates with a higher concentration of shallow defect states while a red-shift corresponds to a greater concentration of deep trap states in these materials prepared using different zinc salts.

The PL behavior and catalytic activity of the ZnS NPs can be analyzed via the corresponding electronic structures. It is well-known that, after photoexcitation, the lattice configuration changes and distorted clusters appear in semiconductor materials allowing electrons to become trapped [69].

In this case, the high-energy photons can excite defect states of electrons localized at different energy levels within the band gap of the material. Therefore, structural defects in the bulk, surface, and/or interface may be trap sites, and electron–hole recombination frequently occurs at these trap sites. The higher surface-to-volume ratios of the ZnS NPs lead to a corresponding enhancement of surface-related emission in the overall luminescence [72,73].

There is a common perception that the predominant structural and electronic disorders in ZnS NPs are sulfur vacancies; it is considered that these are strongly dependent on the morphology, size, structure, and surface composition of the semiconductor materials and significantly affect the physical and chemical behavior of these systems [4]. The ratio of the concentrations of shallow and deep defects observed in the PL measurements can be estimated indirectly. The relationship between shallow and deep defects,  $\rho$ , is an important parameter for understanding the physical and chemical properties of semiconductor materials and can be estimated from the ratio of the left and right areas of the PL emission maximum. Thus, the relationship between shallow and deep defects can be defined as follows:

$$\rho = [\text{shallow defects}]/[\text{deep defects}] \quad (1)$$

Table 1 shows the estimated  $\rho$  values obtained for the ZnS samples prepared using the MAS method. From this model, an increase in the  $\rho$  value is related to a higher

concentration of shallow defects, while a decrease in its value corresponds to a higher concentration of deep defects within the band gap. These results reveal that crystalline parameters may be obtained from the detailed analysis of the PL behavior, and hence can be used to build a close relationship between structure and properties of the ZnS NPs. This finding is very interesting for the monitoring of structural changes as well as their potential applications at the nanoscale. Hence, in this direction, the PL emission behavior may enable the development and the discovery of the new catalysts. Furthermore, the slope observed in the DRS spectra and blue-shift of the PL emission for the ZnS samples indicate varied crystallization of the samples synthesized using different precursors. Obviously, these zinc salts present different bonding forces and the crystalline structure strongly depends on the precursor used, which acts as a surface modifier causing different structural polarization effects on these NPs at short-, medium- and long-ranges, which can be extended to other complex systems. Based on these results, the binding strength of these precursors decreases as follows: ZN < ZC < ZA. This fact provides an excellent opportunity for structural control of these complex functional materials and allows highlighting the crucial role of the precursor as a surface modifier. To further elucidate the structural and electronic properties of these NPs, theoretical calculations were performed.

### Table 1.

The photocatalytic activities of the WZ ZnS NPs prepared using the MAS method with different precursors were evaluated via photodegradation of RhB dye in a 10 mgL<sup>-1</sup> aqueous solution under UV light irradiation. There was no obvious photocatalytic degradation of RhB in the absence of photocatalyst (see Figure 3 (a)).

The results showed that approximately 4% solution discoloration occurred at 210 minutes for all materials, and equilibrium was reached at 15 minutes of reaction. Thus, in all photocatalytic experiments, the dye solution was kept in the dark for 15 minutes under stirring. Remarkably, the experiment to verify the photolysis of RhB in the absence of photocatalyst showed that the RhB hardly decomposed under UV radiation even after 210 min (see Table 2). In contrast, the ZnS samples present good photocatalytic activities, resulting in a significant reduction in the maximum absorption spectra of RhB (554 nm) during the photodegradation process (see Figures 3 (b–d)). **In order to obtain a more appropriate comparison to the well-known and widely applied standards, the synthesized materials were compared with commercial TiO<sub>2</sub> anatase (Aldrich, 99.7%) activity, as shown in Figure 3 (e–f). These results are consistent with other studies [74–76], and hence they reveal a higher photocatalytic activity for ZnS NPs prepared by the MAS method when compared with anatase TiO<sub>2</sub> crystals.** The degradation ratio ( $C/C_0$ ) RhB solution with ZnS NPs and without catalyst versus time and the linear relationship of  $[-\ln(C/C_0)]$  with irradiation time are shown in Figures 3 (e–f).

**Figure 3.**

The photodegradation reaction kinetics of RhB follow a pseudo-first order rate equation. The values of the rate constant can be obtained from the slope of the  $[-\ln(C_0/C_t)] = k_{app}t$ , where  $C_0$  and  $C_t$  are concentrations of RhB at time zero and  $t$ , respectively, and  $k_{app}$  is the apparent first-order rate constant [77]. Correlation coefficients listed in Table 2 ( $R^2 > 0.90$ ) demonstrated a linear relationship between  $[-$

$\ln(C_0/C_t)$ ] and reaction time. Thus, the degradation reaction in the RhB solution catalyzed by ZnS NPs occurs according to model pseudo-first kinetic order.

The ZnS NPs synthesized with ZN exhibited a more superior performance than the samples synthesized with ZC and ZA (see Table 2): 97% of the RhB molecules degraded in the presence of ZnS synthesized with ZN after 210 min of exposure to UV radiation, while the samples synthesized with ZC and ZA degraded 82% and 56% of the RhB molecules, respectively. Clearly, the influence of the precursors as surface modifier of these NPs prepared by the MAS method has played a decisive role in the photocatalytic performance enhancement and allows tailored surface chemistry of ZnS NPs. Through the different precursors used it is possible to observe an influence on the exposed surface facet of the ZnS NPs prepared by the MAS method, leading to acquisition of a non-uniform surface composition (confirmed via XPS analysis), and hence the formation and migration rate of the electron-hole pairs of such particles as well as the specific surface area of these catalysts affecting the heterogeneous processes. Although an increase occurs in specific surface area of the ZnS NPs governed by different precursors and our results reveal how the reaction rate of the photocatalytic process decreases in this sequence:  $ZN > ZC > ZA$ . In the photocatalytic experiments, a heterogeneous process takes place on the surface of the photocatalyst and then an analysis of the photocatalytic activity of these materials was obtained using the ratio of the rate constant and the specific surface area of the material, instead of using only the degradation percentage or the reaction constant ( $k$ ). The relationship between the values of the reaction constant with specific surface area ( $k/SSA$ ) is shown in Table 2.

**Table 2.**

Based on the obtained results, we can show that up to 150 min of reaction the photocatalytic behavior of the both ZN and ZC samples is quite similar, and hence are far superior to present by the ZA sample. From the 150 min onward, the photocatalytic activity of the ZC and ZA samples becomes stabilized, whereas for the ZN sample this stabilization occurs only at 210 min.

The redox ability of ZnS can be obtained by determining the energy positions of both conduction band (CB) and valence band (VB) edges. Generally, the more positive the potentials of the VB are, the stronger the holes of oxidation capacity are while the more negative the potentials of the CB are, the stronger the electrons of reduction ability are [78]. According with the band model, in principle, the VB and CB edges positions can be calculated according to the equation as follows [79]:

$$E_{\text{VB}} = \chi - E^e + 0.5 E_g \quad (2)$$

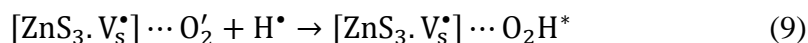
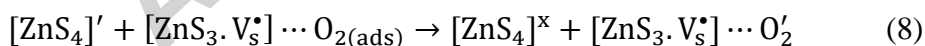
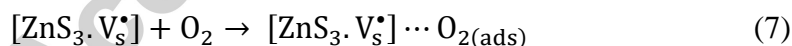
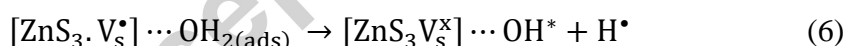
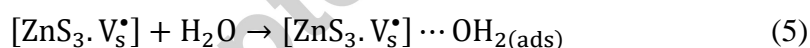
$$E_{\text{CB}} = E_{\text{VB}} - E_g \quad (3)$$

where  $E_{\text{VB}}$  and  $E_{\text{CB}}$  are the VB and CB edges potentials, respectively. The  $\chi$  being the Mulliken electronegativity, which is calculated as the arithmetic mean of the atomic electron affinity and the first ionization energy of the constituent atoms in the semiconductor [80,81]. So in principle, the  $\chi$  is calculated to be about 5.26 eV for the ZnS [78,81].  $E^e$  is the energy of free electrons on the hydrogen scale ( $\sim 4.5$  eV) and  $E_g$  is the band gap energy values. Figure 4 shows the schematic diagrams (V versus NHE/eV) for ZnS NPs as photocatalyst. In particular, the VB maximum potential for as-prepared ZnS illustrate a variation in the values from 2.485 eV for ZN to 2.520 eV for ZC and ZA, respectively, which in all cases is more positive than  $\text{O}_2/\text{H}_2\text{O}$  (1.23 V),  $\text{H}_2\text{O}_2/\text{H}_2\text{O}$  (1.77 V) and  $\text{O}_3/\text{H}_2\text{O}$  (2.07 V) [82]. Therefore, our results reveals that the

ZnS as-prepared has the ability to oxidize H<sub>2</sub>O to produce O<sub>2</sub> and decompose many organic pollutants. The CB minimum potential of as-prepared ZnS illustrate a variation in the values is about -0.965 eV for ZN, -1.020 eV for ZC and -1.000 eV for ZA, which is more negative than H<sup>+</sup>/H<sub>2</sub> (0 V), and hence can reduce H<sup>+</sup> to H<sub>2</sub>, which shows that ZnS can split water into H<sub>2</sub>. These results calculated for the top of the VB and the bottom of the CB for the as-prepared ZnS NPs are in good agreement with literature values [78,81], and hence reveals a high redox ability for such samples.

**Figure 4.**

We propose a model to explain the relevant elementary steps involved in the adsorbed and dissociation process of the H<sub>2</sub>O and O<sub>2</sub> species on the ZnS surface, in accordance with the complex clusters based on the Kröger–Vink notation, and can be summarized by the following equations:



We note that the structural defects on the crystal surface play an important role as the active surface sites responsible for dissociation of adsorbed H<sub>2</sub>O and O<sub>2</sub> species and lead to the production of OH<sup>\*</sup> and O<sub>2</sub>H<sup>\*</sup> radicals, which are the most strongly



oxidizing species in this process. Through a combination of scanning tunneling microscopy (STM) experiments and DFT calculations, Schaub *et al.* [83] reported the unambiguous determination of the active surface sites responsible for dissociation of water molecules adsorbed on the (110) surface of rutile TiO<sub>2</sub>, i.e., water dissociation was only energetically feasible at the defect sites in the (110) surface of the TiO<sub>2</sub>. On the other hand, Li *et al.* [84] presented a systematic study of the facet-dependent photo-electrochemical performance of anatase TiO<sub>2</sub> nanomaterials with predominantly exposed (101), (010), and (001) TiO<sub>2</sub> surfaces; their results were based on a combination of photocatalysis measurements, characterization of photo-electrochemical performance in dye-sensitized solar cell devices, time-resolved spectroscopy, and first-principles calculations. This suggests that the photochemical performance highly correlates with the surface energy and number of under-coordinated surface atoms, and hence facilitate the formation rate of the OH\* and O<sub>2</sub>H\* radicals. **In addition, Dong et al. [85] have reported theoretical calculations on the distorted tetrahedral [ZnS<sub>4</sub>] clusters in WZ ZnS systems could in principle be beneficial to the separation process of the photogenerated charge carriers. These theoretical results are in very good agreement with our experimental findings.**

At this stage, however, it is important to note that direct comparison of theoretical and experimental results is essential and necessary to provide in depth understanding of the electronic and structural properties of NPs and their dependence on the synthetic route and environmental growth conditions. Theoretical calculations can be used not only for interpretation of experiments, but also for prediction of important aspects of new properties that emerge at the nanoscale [51,86,87]. The morphology, surface DOS, and surface Fermi level position (and the reliability of the theory to predict these) is key for explaining the physical and chemical behavior at the molecular level, because there is a

high surface-to-bulk ratio and these effects are capable of governing the optical, electronic, and photocatalytic properties [86,87]. As far as we know, to date, systematic studies towards revealing the surface-structure and facet-dependent catalytic activity relationships of the ZnS nanocrystals are still very lacking, and hence remain a fascinating challenge to theoretical and experimental researchers. Therefore, theoretical simulations were performed at the DFT level to gain a better understanding of the experimental data, to distinguish between the bulk and surface contributions to the DOS, and to resolve the structural and electronic effects on the physical and chemical properties of the WZ ZnS NPs prepared using the MAS method with different precursors.

Furthermore, it is expected that the specific surface area of ZnS NPs plays a key role during photocatalytic degradation of RhB dye molecules; therefore, both the surface atomic structure and electronic band structure were systematically investigated via DFT calculations. This has enabled the understanding of the atomic and electronic properties of a crystal surface, which has provided some insight into the features of single crystal facets relevant to subsequent technological applications [88-90]. Table 3 shows that the calculated surface energy values for the (0001), (10 $\bar{1}$ 0), (11 $\bar{2}$ 0), (10 $\bar{1}$ 1), and (10 $\bar{1}$ 2) facets of ZnS are 1.40, 0.56, 0.41, 1.11, and 1.24 Jm<sup>-2</sup>, respectively. The (11 $\bar{2}$ 0) surface is the most stable facet and the surface stability order agrees with that shown in other theoretical results [52,91]. Different ZnS surfaces have different catalytic activities, with (0001) being the most reactive surface. An analysis of the results reported in Table 3 reveals a relationship between the energy gap and surface energy. The lower value of the band gap arises from the main surface reactivity of the exposed face during catalytic processes, which provides a valuable tool for interpretation of the experimental results. The results of the analysis of the projected total DOS of the (0001), (10 $\bar{1}$ 0), (10 $\bar{1}$ 1), and

(11 $\bar{2}$ 0) facets of ZnS are shown in Figures 5 (a–d).

**Figure 5.**

Insights from theoretical calculations and experimental evidence can be related to develop a simple model that can be used to better understand crystal facet modulation with the aim of providing new insights into the structural design of the NPs as well as optical and photocatalytic mechanisms at the atomic level. The Wulff construction is a well-recognized method for determining the equilibrium morphology of complex functional materials [43-50,91,92], providing valuable thermodynamic information about frequent crystalline surface energies and constructing the most stable crystal shape.

Figure 6(a) shows the predicted crystal morphology based on the Wulff construction derived from the calculated surface energy values reported in Table 3. In particular, the Visualization for Electronic and Structural Analysis (VESTA) program (version 3.1.8 for Windows) [93] was used for the Wulff construction as well as in the visualization of the morphological mapping of ZnS crystals.

**Table 3.**

Indeed, the TEM results reported in the Supporting Information show that the ZnS samples have an irregular spherical-like shape. Theoretical results predict a cylindrical-like morphology for WZ ZnS with a large non-polar area exposing the (11 $\bar{2}$ 0) facets and (0001) and (10 $\bar{1}$  $\bar{2}$ ) facets at the top and bottom of the cylindrical-like morphology of ZnS nanocrystals (see Figure 6(a)), which is in agreement with other reported results [91].

It is important to note that present theoretical results consider the crystals to be *in*

*vacuum*. The exposed surface planes during the crystal growth process (in particular, those crystal facets with large surface energies) will shrink or even diminish rapidly during the crystal growth process as a result of minimizing the surface energy of small crystals [94], these eventually disappear or have decreased participation in the generated morphology. In this way, the surface is usually encircled by planes with lower surface energies, resulting in a slower crystal growth rate [95]. In this sense, the achieved morphology can be modulated by changing the surface energies of the different exposed surface planes during the crystal growth process thereby modifying the nanomaterial properties.

**Figure 6.**

Basically, controlling the crystal growth process is complex and very sensitive to the chemical environment and internal crystal structures mandated by the synthesis strategy. The type of precursor can be selectively adsorbed onto specific exposed surface planes, thus changing the surface energies and inhibiting a particular active site during the crystal growth process. Therefore, the final morphology based on Wulff's rule will also change. Several insightful review papers have reported experimental and theoretical calculations on the important aspects that govern crystal shape modulation in semiconductor and metal NPs [96-100].

Therefore, in order to obtain the experimental morphology, a proposed hypothesis is that the interaction between precursors decreases the surface energy of the exposed (0001) surface with respect to those of the (10 $\bar{1}$ 0), (11 $\bar{2}$ 0), (10 $\bar{1}$ 2), and (10 $\bar{1}$ 1) surfaces. For the present case, a morphological simulation is reported in Figures 6(b)–(e), in which the Wulff's crystal of the optimized ZnS and different morphologies are obtained

assuming a different surface energy ratio; this generates a controllable and predictable way to model the morphology in greater depth.

The precursors (i.e., ZN, ZC, or ZA) affect the particle stability and morphology. Theoretically, the achieved morphology depends on the surface energy of each crystal termination. An analysis of Figure 6 shows that the ideal morphology is a hexagonal prism. When the ratio of the energy surface is varied by decreasing the value of the (0001) surface energy through some perturbation, the results (shown in Figure 5(b)) reveal that the  $(10\bar{1}2)$  facet is disfavoured and the  $(10\bar{1}1)$  facet is favoured instead. The  $(10\bar{1}1)$  surface appears to be favoured with increasing  $(10\bar{1}2)$  surface energy (shown in Figure 6(c)). When the ratio of the energy surface is varied by decreasing or increasing the  $(10\bar{1}0)$  or  $(11\bar{2}0)$  surface energy, the  $(10\bar{1}0)$  facet is exposed in the morphology (shown in Figure 6(d)). In particular, a more or less spherical-like morphology can be achieved by modifying the  $(10\bar{1}2)$  energy surface, which is in agreement with the experimental TEM images. Therefore, it was found that the experimental morphology can be modulated by tuning the (0001) surface stability, which depends on the precursor.

From this result, it is suggested that the samples synthesized with ZN present a greater percentage of exposed (0001) surface than the samples synthesized with ZC and ZA, as revealed by the higher photocatalytic activity of the former NPs. This fact has important consequences for understanding the photocatalytic mechanism of these NPs, as corroborated by the recently reported study of the facet-dependent photocatalytic performance of ZnO nanostructures [101], which showed the following order for the photodegradation efficiencies of the  $(0001) > (10\bar{1}1) > (10\bar{1}0)$  surfaces [81,90].

This study shows that the different types of defects (i.e., shallow and deep) in ZnS NPs are responsible for generating improved catalytic activity, which is strongly

influenced by the nanoscale structural and electronic order-disorder effects. A relationship between the PL behavior,  $\rho$ , and photocatalytic activity is evident. A close similarity between experimental and theoretical morphologies can be achieved by tuning the (0001) surface stability, which depends on the nature of the precursor as a surface modifier. Therefore, this work is important for a better understanding of the optical, structural, and electronic properties and photocatalytic applications of the ZnS NPs. In this direction, we envision that a more realistic theoretical simulation based on representation of one specific type of surface exposed, combined with a highly sensitive detection of in situ and operando experiments under different conditions, in principle, may provide a unique opportunity to confirm our hypothesis that attempts to describe a close relationship between the photocatalytic activities and PL properties, i.e., as function of the type of its exposed surface and its surface composition (i.e., the acquired understandings of its morphologies) of the as-prepared ZnS nanocrystals and is a helpful approach to quantitative measurements at the nanoscale.

## Conclusions

In summary, an easy MAS method employing different precursors (i.e., ZN, ZC, or ZA) has been used for the synthesis of ZnS NPs in order to tune optical properties and photocatalytic applications. The results confirm that the hexagonal WZ phase is obtained at low temperatures with a high reaction yield, thereby providing a strategy for obtaining these materials on a large scale. The PL behavior of the ZnS NPs is precursor-dependent and undergoes a red-shift that is linked to specific defects in the crystal framework with maximum emissions at 478, 505, and 555 nm for the NPs synthesized from ZN, ZC, and ZA, respectively, when excited by a 350.7 nm laser line at room temperature. The experimental results demonstrated that the ZnS NPs had good

photocatalytic activity for degradation of RhB under UV irradiation, which can be attributed to its unique structure. Increased photocatalytic degradation was observed for samples synthesized using different precursors in the following order: ZA<ZC<ZN.

**Our findings reveal the crucial role of structural defects that can in principle provide a more efficiently separation of the photogenerated charges.**

First principles calculations complement experimental work and we have contributed to increasing the knowledge on the structural, chemical and electronic changes of ZnS and we propose that the samples synthesized with ZN present a greater percentage of exposed (0001) surface than those synthesized with ZC and ZA. Overall, our investigation clearly reveals that theoretical simulations constitute a helpful strategy to explore experimental conditions and the present study demonstrates that the theoretical findings are in good agreement with the experimental data, and show a connection between the PL behavior and photocatalytic activity of ZnS NPs.

Overall, our investigation clearly reveals that theoretical simulations constitute a helpful strategy to explore experimental conditions. The present synthetic strategy should be a promising fabrication technique for a simple and rapid production of the photocatalysts for the efficient removal of pollutants and open new avenues toward rational design that will benefit the development of novel complex functional materials with desired features and functions. In short, we believe that the focus on factors that control the physical and chemical properties of NPs provides clues that will help contribute to meeting the great demands for environmental sustainability in the near future.

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## References

- [1] Z. Zhang, J. She, H. Chen, S. Deng, N. Xu, *J. Mater. Chem. C*, 1 (2013) 4970-4978.
- [2] X. Fang, T. Zhai, U. K. Gautam, L. Li, L. Wu, Y. Bando, D. Golberg, *Prog. Mater.Sci.*, 56 (2011) 175-287.
- [3] S. Shen, Q. Wang, *Chem. Mater.*, 25 (2013) 1166–1178.
- [4] F. A. La Porta, J. Andrés, M. S. Li, J. R. Sambrano, J. A. Varela, E. Longo, *Phys. Chem. Chem. Phys.*, 16 (2014) 20127-20137.
- [5] C. S. Tiwary, S. Saha, P. Kumbhakar, K. Chattopadhyay, *Cryst. Growth Des.*, 14 (2014) 4240–4246.
- [6] B. Liu, Q. Liu, C. Tong, *Colloids Surface A*, 434 (2013) 213-219.
- [7] Y. Zhu, Q. Ruan, F. Xu, *Nano. Res.*, 2 (2009) 688-694.
- [8] A. Kumar, R. K. Upadhyay, *J Mater Sci: Mater Electron.*, 26 (2015) 2430-2435.
- [9] F. A. La Porta, M. M. Ferrer, Y. V. B. Santana, C. W. Raubach, V. M. Longo, J. R. Sambrano, E. Longo, J. Andrés, M. S. Li, J. A. Varela, *J. Alloys Compd.*, 556 (2013) 153-159.
- [10] N. Srinivasan, S. Thirumaran, *C. R. Chimie.*, 17 (2014) 964–970.



- [11] H. Zhang, B. R. Hyun, F. W. Wise, R. D. Robinson, *Nano Lett.* 12 (2012) 5856-5860.
- [12] Z. Q. Mamiyev, N. O. Balayeva, *Chem. Phys. Lett.*, 646 (2016) 69-74.
- [13] S. Ummartyotin, Y. Infahsaeng, *Renew. Sustainable Energy Rev.*, 55 (2016) 17-24.
- [14] N. H. Abdullah, Z. Zainal, S. Silong, M. I. M. Tahir, K. B. Tan, S. K. Chang, *Mater Chem Phys.*, 176 (2016) 33-41.
- [15] K. Deka, M. P. C. Kalita, *Braz J Phys.*, 45 (2015) 36-40.
- [16] L. Yin, D. Zhang, D. Wang, X. Kong, J. Huang, F. Wang, Y. Wu, *Mater. Sci. Eng. B*, 208 (2016) 15-21.
- [17] P. Hu, G. Gong, F. Zhan, Y. Zhang, R. Li and Y. Cao, *Dalton Trans.*, 45, (2016) 2409-2416.
- [18] R. Ramachandran, M. Saranya, P. Kollu, B. P. C. Raghupathy, S. K. Jeong, A. N. Grace, *Eletrochim Acta.*, 178, (2015) 647-657.
- [19] P. Hu, G. Gong, F. Zhan, Y. Zhang, R. Li, Y. Cao, *Dalton Trans.*, 45 (2016) 2409-2416.
- [20] K. Vijaianand, M. Karlchinnu, R. Mohankumar, R. Mohan and R. Jayavel, *Appl. Surf. Sci.* 255 (2009) 8879-8882.
- [21] X. Yan, E. Michael, S. Komarnenic, J. R. Brownsond and Z. F. Yana, *Ceram. Int.* 39 (2013) 4757-4763
- [22] **L. Wang, P. Wang, B. Huang, X. Ma, G. Wang, Y. Dai, X. Zhang and X. Qin, *Appl. Surf. Sci.* (2016) in press (DOI: 10.1016/j.apsusc.2016.06.159)**

- [23] **R. Wang, H. Liang, J. Hong and Z. Wang, J. Photochem. Photobiol., A 325 (2016) 62-67**
- [24] M. M. Ferrer, Y. V. B. de Santana, C. W. Raubach, F. A. La Porta, A. F. Gouveia, E. Longo, J. R. Sambrano, *J. Mol. Model.*, 20 (2014) 2375-2383.
- [25] Y. V. B. Santana, C. W. Raubach, M. M. Ferrer, F. A. La Porta, J. R. Sambrano, V. M. Longo, E. R. Leite, E. Longo, *J. Appl. Phys.*, 110 (2011) 123507.
- [26] **L. L. Wang, C. Z. Zhao, L. P. Kang, D. W. Liu, H. C. Zhao, S. P. Hao, Y. K. Zhang, Z. P. Chen and X. J. Li, Appl. Surf. Sci. 384 (2016) 530-533.**
- [27] **S. Kaur, S. Sharma and S. K. Kansal, Superlattices Microstruct. 98 (2016) 86-95.**
- [28] **S. Ahluwalia, N. T. Prakash, R. Prakash and B. Pal, Chem. Eng. J. 306 (2016) 1041-1048.**
- [29] **K. Kalpana and V. Selvaraj, J. Ind. Eng. Chem. 41 (2016) 105-113.**
- [30] R. Bose, U. Thupakula, J. K. Bal, N. Pradhan, *J. Phys. Chem. C*, 116 (2012) 16680–16686.
- [31] J. G. Zhao, H. H. Zhang, *Superlattices Microstruct.*, 51 (2012) 663–667.
- [32] S. Li, Z. Wu, W. Li, Y. Lin, R. Zhuo, D. Yan, W. Jun, P. Yan, *CrystEngComm*, 15 (2013) 1571-1577.
- [33] M. S. Niasari, F. Davar, H. Seyghalkar, E. Esmaeili, N. Mir, *CrystEngComm*, 13 (2014) 2948–2954.
- [34] M. Scocioreanu, L. Mihut, M. Baibarac, I. Baltog, *Phys. Status Solidi B*, 250 (2013) 1426-1431.
- [35] Y. Cheng, Z. Lin, H. Lu, L. Zhang, B. Yang, *Nanotechnology*, 25 (2014) 115601.
- [36] F. Kurnia, Y. Hau Ng, Y. Tang, R. Amal, N. Valanoor, J. N. Hart, *Cryst. Growth Des.*, 16 (2016) 2461–2465.

- [37] C. W. Raubach, Y. V. B. de Santana, M. M. Ferrer, P. G. C. Buzolin, J. R. Sambrano, E. Longo. *Dalton Trans.*, 42 (2013) 11111-11116.
- [38] S. Xiong, B. Xi, C. Wang, D. Xu, X. Feng, Z. Zhu, Y. Qian, *Adv. Funct. Mater.*, 17 (2007) 2728-2738.
- [39] X. Yu, J. Yu, B. Cheng, B. Huang, *Chem. Eur. J.*, 15 (2009) 6731 – 6739.
- [40] H. S. Park, W. B. Ko, *J. Nanosci. Nanotechnol.*, 14 (2014) 8646-8653.
- [41] H. Naeimi, H. Foroughi, *New J. Chem.*, 39 (2015) 1228-1236.
- [42] S. Sarkar, A. K. Guria, N. Pradhan, *Chem. Commun.*, 49 (2013) 6018-6020.
- [43] M. R. D. Bomio, R. L. Tranquilin, F. V. Motta, C. A. Paskocimas, R. M. Nascimento, L. Gracia, J. Andres, E. Longo, *J. Phys. Chem. C*, 117 (2013) 21382-21395.
- [44] R. A. Roca, J. C. Sczancoski, I. C. Nogueira, M. T. Fabbro, H. C. Alves, L. Gracia, L. P. S. Santos, C. P. de Sousa, J. Andrés, G. E. Luz Jr, E. Longo, L.S. Cavalcante, *Catalysis Science & Technology*, 5 (2015) 4091-4107.
- [45] R. A. Roca, P. S. Lemos, J. Andres, Longo, E. *Chem. Phys. Lett.* 590 (2016), 644, 68–72.
- [46] M. T. Fabbro, C. Saliby, L. R. Rios, F. La Porta, L. Gracia, M. Li, J. Andrés, L. P. Serejo dos Santos, E. Longo, *Sci. Technol. Adv. Mater*, 16 (2015) 065002.
- [47] R. C. de Oliveira, M. Assis, M. M. Teixeira, M. D. P. da Silva, M. S. Li, J. Andres, L. Gracia, E. Longo, *J. Phys. Chem. C*, 120 (2016) 12254-12264
- [48] M. T. Fabbro, L. Gracia, G. S. Silva, L. P. S. Santos, J. Andrés, E. Cordoncillo, E. Longo, *J. Solid State Chem.* 239, (2016) 220-227.
- [49] G. Botelho, J. Andres, L. Gracia, L. S. Matos, E. Longo, *Chem PhusChem.*, 81 (2016) 202-212 .

- [50] G. Botelho, J.C. Sczancoski, J. Andrés, L. Gracia and E. Longo, *J. Phys. Chem. C*, 119 (2015) 6293-6306.
- [51] F. A. La Porta, L. Gracia, J. Andrés, J. R. Sambrano, J. A. Varela, E. Longo, *J. Am. Ceram. Soc.*, 97 (2014) 4011-4018.
- [52] C. A. Feigl, A. S. Barnard, S. P. Russo, *Phys. Chem. Chem. Phys.*, 14 (2012) 9871–9879.
- [53] X. Meng, H. Xiao, X. Wen, W. Goddard III A, S. Li, G. Qin, *Phys. Chem. Chem. Phys.*, 15 (2013) 9531-9539.
- [54] E. Balantseva, G. Berlier, B. Camino, M. Lessio, A. M. Ferrari, *Study, J. Phys. Chem. C*, 118 (2014) 23853-23862.
- [55] J. M. Azpiroz, E. Mosconi, F. D. Angelis, *J. Phys. Chem. C*, 115 (2011) 25219-25226.
- [56] J. M. Azpiroz, E. Mosconi, J. M. Ugalde, F. D. Angelis, *J. Phys. Chem. C*, 118 (2014) 3274-3284.
- [57] H. Zhang, B. Gilbert, F. Hung, J. F. Banfield, *Nature*, 424 (2003) 1025-1029.
- [58] G. Kresse, G. J. Hafner, *Phys. Rev. B*, 49 (1994) 14251.
- [59] G. Kresse, G. J. Furthmuller, *Comput. Mater.Sci.*,6 (1996) 15-50.
- [60] J. P. Perdew, K. Burke, M. Ernzerhof, *Phys. Rev. Lett.*, 77 (1997) 3865.
- [61] G. Kresse, D. Joubert, *Phys. Rev. B*, 59 (1999) 1758.
- [62] H. J. Monkhorst, J. D. Pack, *Phys. Rev. B: Solid State*, 13 (1976) 5188.
- [63] J. F. Xu, W. Ji, J. Y. Lin, S. H. Tang, Y. W. Du, *Appl. Phys. A*, 66 (1998) 639–641.
- [64] C. D. Wagner, G. E. Muilenberg, *Handbook of X-ray photoelectron spectroscopy: a reference book of standard data for use in x-ray photoelectron spectroscopy*, Physical Electronics Division, Perkin-Elmer Corp., 1979.

- [65] Y -P. Zhu, J. Li, T -Y. Ma, Y -P. Liu, G. Du, Z -Y. Yuan, *J. Mater. Chem. A*, 2 (2014) 1093-1101.
- [66] V. Etacheri, M. K. Seery, S. J. Hinder, S. C. Pillai, *Adv. Funct. Mater.*, 21 (2011) 3744-3752.
- [67] V. Kumar, V. Kumar, S. Som, J. H. Neethling, M. Lee, O. M. Ntwaeaborwa, H. C. Swart, *Nanotechnology*, 25 (2014) 135701-9.
- [68] J -S. Hu, L -L. Ren, Y. Guo, H -P. Liang, A -M. Cao, L -J. Wan, C -L. Bai, *Angew. Chem.* 117 (2005) 1295 -1299.
- [69] F. A. La Porta, J. Andrés, M. V. G. Vismara, C. F. O. Graeff, J. R. Sambrano, M. S. Li, J. A. Varela, E. Longo, *J. Mater. Chem. C*, 2 (2014) 10164-10174.
- [70] J. M. Ziman, *Models of Disorder: The Theoretical Physics of Homogeneously Disordered Systems*, Cambridge University Press, 1979.
- [71] S. K. Estreicher, T. M. Gibbons, M. B. Bebek, *J. Appl. Phys.*, 117 (2015) 112801.
- [72] Y. Gong, T. Andelman, G. F. Neumark, S. O'Brien, I. L. Kukovsky. *Nanoscale Res. Lett.*, 2 (2007) 297-302.
- [73] E. Rauwel, A. Galeckas, P. Rauwel, M. F. Sunding, H . Fjellvag, *J. Phys. Chem. C*, 115 (2011) 25227-25233.
- [74] **U.G. Akpan, B.H. Hameed. *J. Hazard. Mater.* 170 (2009) 520.**
- [75] **G. Plantard, T. Janin, V. Goetz, S. Brosillon. *Appl. Catal., B.* 115-116 (2012) 38-44.**
- [76] **D. Yang, H. Liu, Z. Zheng, Y. Yuan, J.-C. Zhao, E.R. Waclawik, X. Ke, H. Zhu. *J. Am. Chem. Soc.* 131 (2009) 17885-17893**
- [77] M. Huang, S. Weng, B. Wang, J. Hu, X. Fu, P. Liu, *J. Phys. Chem. C*, 118 (2014) 25434-25440.
- [78] X. Chen, S. Shen, L. Guo, S.S. Mao, *Chem. Rev.* 110 (2010) 6503-6570.

- [79] Y. Xu, M. A. A. Schoonen, *Am. Mineral.*, 85 (2000) 543-556.
- [80] H. Shi, Z. Li, J. Kou, J. Ye, Z. Zou, *J. Phys. Chem. C* 115 (2010) 145-151.
- [81] C. Mondal, A. Singh, R. Sahoo, A. K. Sasmal, Y. Negishi, T. Pal, *New J.Chem.*, 39 (2015) 5628-5635.
- [82] J. Tang, J. Ye, *Chem. Phys. Lett.* 410 (2005) 104
- [83] R. Schaub, P. Thostrup, N. Lopez, E. Laegsgaard, I. Stensgaard, J. K. Nørskov, F. Besenbacher, *Phys. Rev. Lett.*, 87 (2001) 266104.
- [84] C. Li, C. Koenigsmann, W. Ding, B. Rudshiteyn, K. R. Yang, K. P. Regan, S. J. Konezny, V. S. Batista, G. W. Brudvig, C. A. Schmuttenmaer, J -H. Kim, *J. Am. Chem. Soc.*, 137 (2015) 1520-1529.
- [85] M. Dong, J. Zhang and J. Yu, *APL Mater.* 3 (2015) 104404.**
- [86] J. K. Nørskov, T. Bligaard, J. Rossmeisl, C. H. Christensen, *Nat. Chem.*, 1 (2009) 37-46.
- [87] M. Hjort, S. Lehmann, J. Knutsson, A. A. Zakharov, Y. A. Du, S. Sakong, R. Timm, G. Nylund, E. Lundgren, P. Kratzer, K. A. Dick, A. Mikkelsen, *ACS Nano*, 8 (2014) 12346-12355.
- [88] J. Andrés, L. Gracia, A. F. Gouveia, M. M. Ferrer, E. Longo, *Nanotechnology*, 2015, 26, 405703.
- [89] M. M. Ferrer, A. F. Gouveia, L. Gracia, E. Longo, J. Andrés, *Modelling Simul. Mater. Sci. Eng.* 2016, 24, 025007.
- [90] M. C. Oliveira, L. Gracia, I. C. Nogueira, M. F. C. Gurgel, J. M. R. Mercury, E. Longo, J. Andrés, *Ceram. Inter.* 2016, 42, 10913
- [91] S. Hamad, S. Cristol, C. R. A. Catlow, *J. Phys. Chem. B*, 106 (2002) 11002–11008.
- [92] A. S. Barnard, C. A. Feigland, S. P. Russo, *Nanoscale*, 2 (2010) 2294–2301.
- [93] K. Momma, F. Izumi, *J. Appl. Crystallogr.*, 44 (2011) 1272.

- [94] C -F. Huo, B -S. Wu, P. Gao, Y. Yang, Y -W. Li, H. Jiao, *Angew. Chem.*, 123 (2011) 7541-7544.
- [95] C. Burda, X. B. Chen, R. Narayanan, M. A. El-Sayed, *Chem. Rev.*, 105 (2005) 1025-1102.
- [96] M. A. Lovette, A. R. Browning, D. W. Griffin, J. P. Sizemore, R. C. Snyder, M. F. Doherty, *Ind. Eng. Chem. Res.*, 47( 2008) 9812-9833.
- [97] A. Seyed-Razavi, I. K. Snook, A. S. Barnard, *J. Mater. Chem.*, 20 (2010) 416-421.
- [98] J. Y. Chen, B. Lim, E. P. Lee, Y. N. Xia, *Nano Today*, 4 (2009) 81.
- [99] Z. Y. Jiang, Q. Kuang, Z. X. Xie, L. S. Zheng, *Adv. Funct. Mater.*, 20 (2010) 3634-3645
- [100] K. Lee, M. Kim, H. Kim, *J. Mater. Chem.*, 20 (2010) 3791-3798.
- [101] Y. V. Kaneti, Z. Zhang, J. Yue, Q. M. D. Zakaria, C. Chen, X. Jiang, A. Yu, *Phys. Chem. Chem. Phys.*, 16 (2014) 11471-11480.

**Figure Captions:**

**Figure 1:** Zinc 2p<sub>3/2</sub> and sulfur 2p XPS spectra of wurtzite ZnS NPs: (a-b) ZN; (c-d) ZC and (e-f) ZA.

**Figure 2:** (a) DRS spectra and (b) PL spectra at room temperature of aggregated crystalline ZnS NPs obtained by a MAS method.

**Figure 3:** UV-vis absorption spectra of solutions in the process of the photodegradation of RhB under UV irradiation, (a) without catalyst and in the presence of ZnS NPs produced from (b) ZN, (c) ZC and (d) ZA as precursors, respectively. (e) Photodegradation curves and (f) First-order kinetic curves for the discoloration of RhB solutions.

**Figure 4:** Schematic diagram for the redox ability of ZnS NPs.

**Figure 5:** Projected total DOS of the a) (0001), b) (10 $\bar{1}$ 0), c) (11 $\bar{2}$ 0) and d) (10 $\bar{1}$ 2) of ZnS surfaces.

**Figure 6:** Morphology map for the ZnS crystal obtained from Wulff's construction and assuming different surface energy values of the different faces.



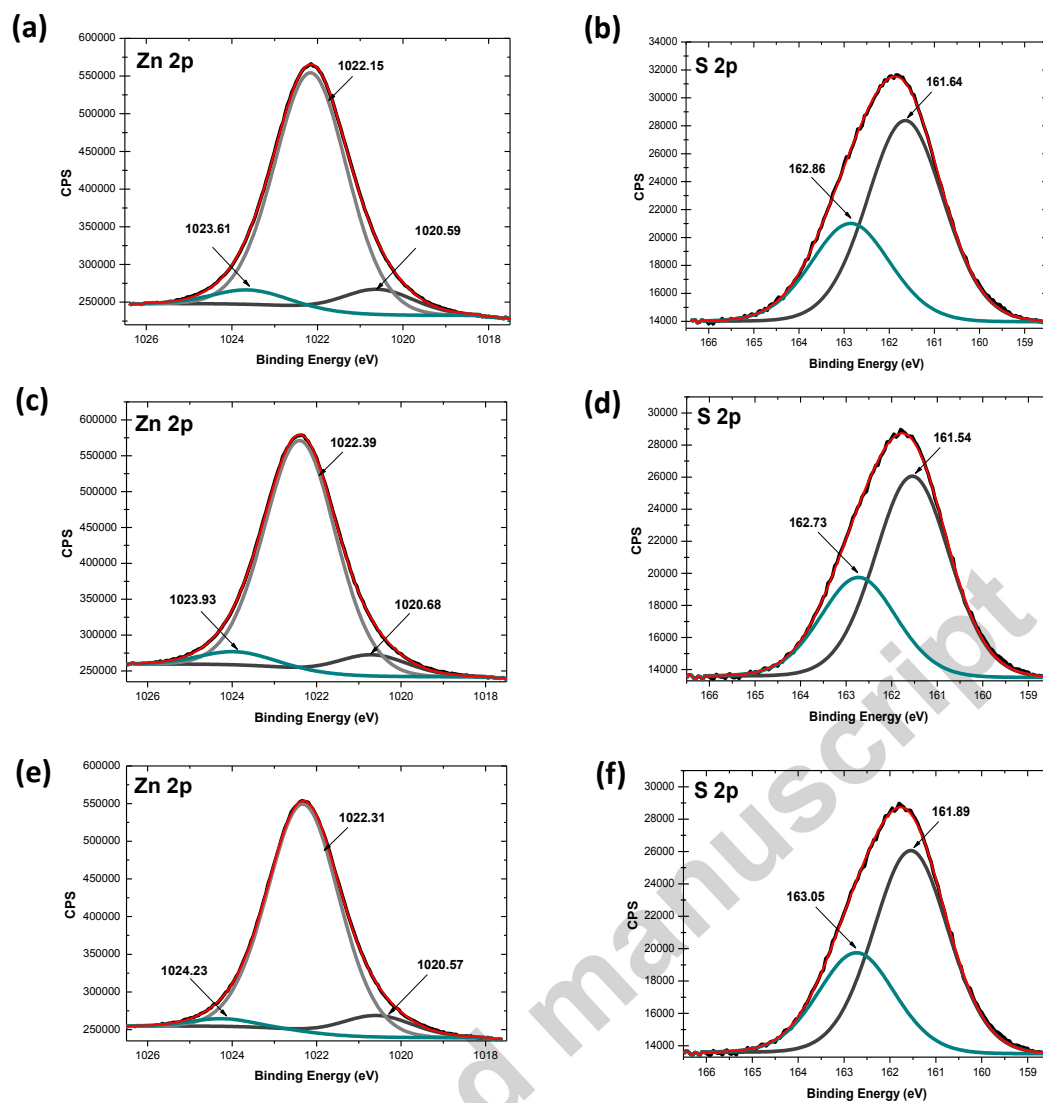


Figure 1.

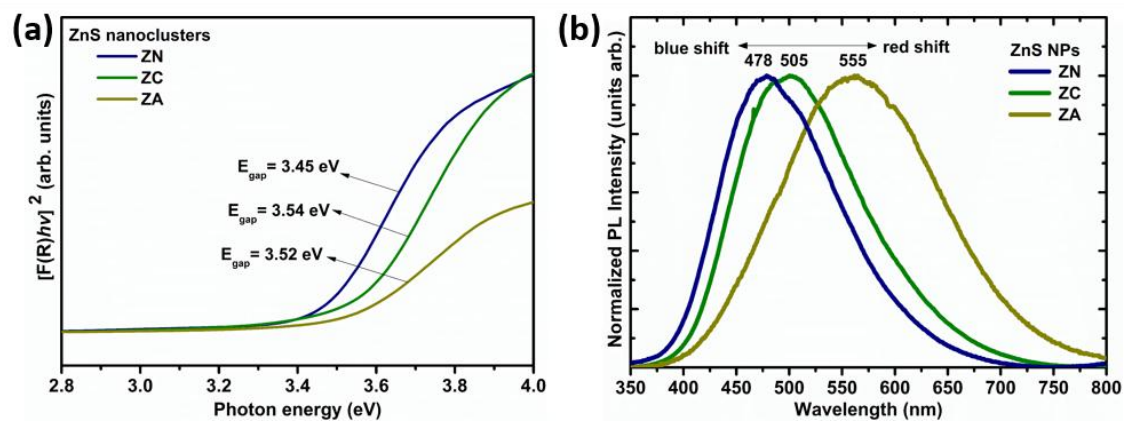


Figure 2.

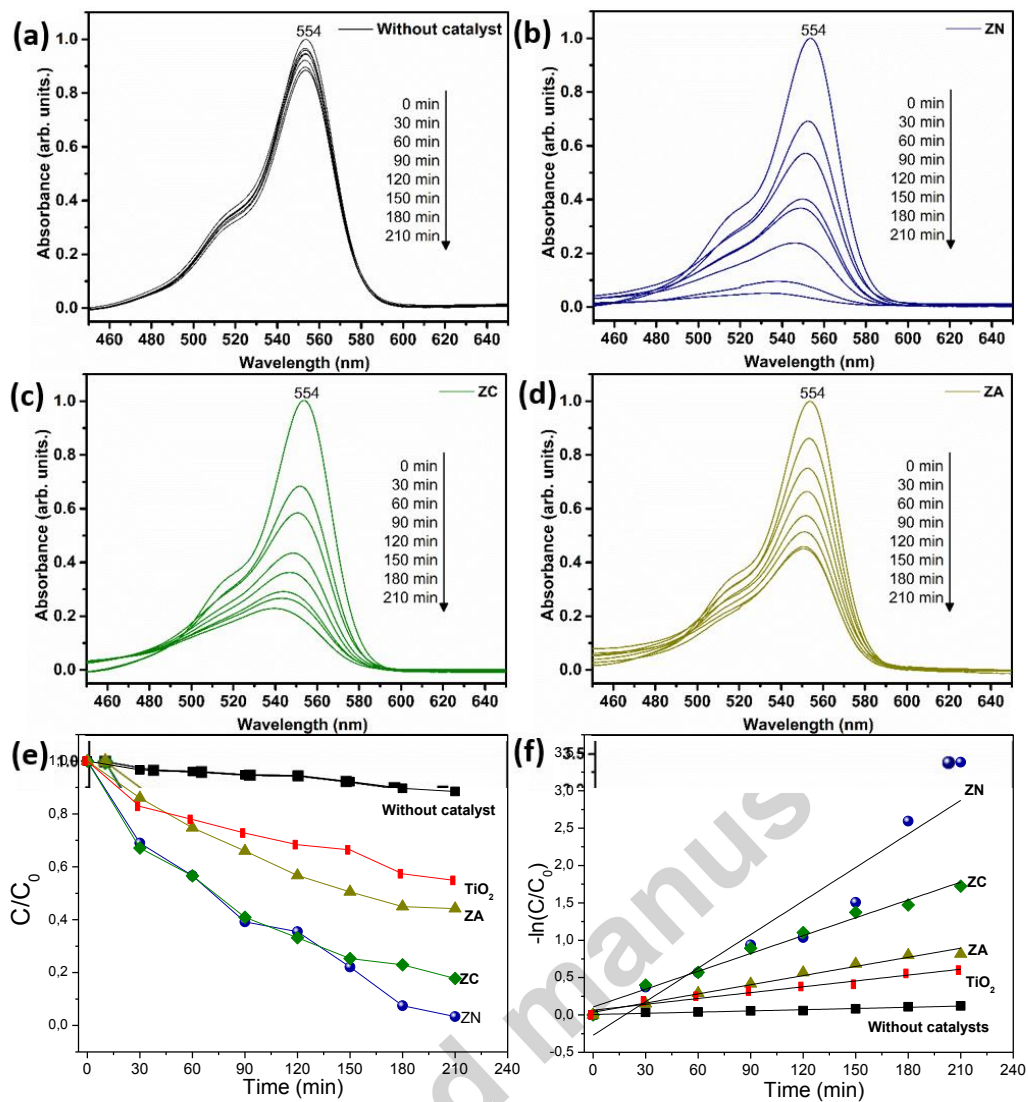


Figure 3.

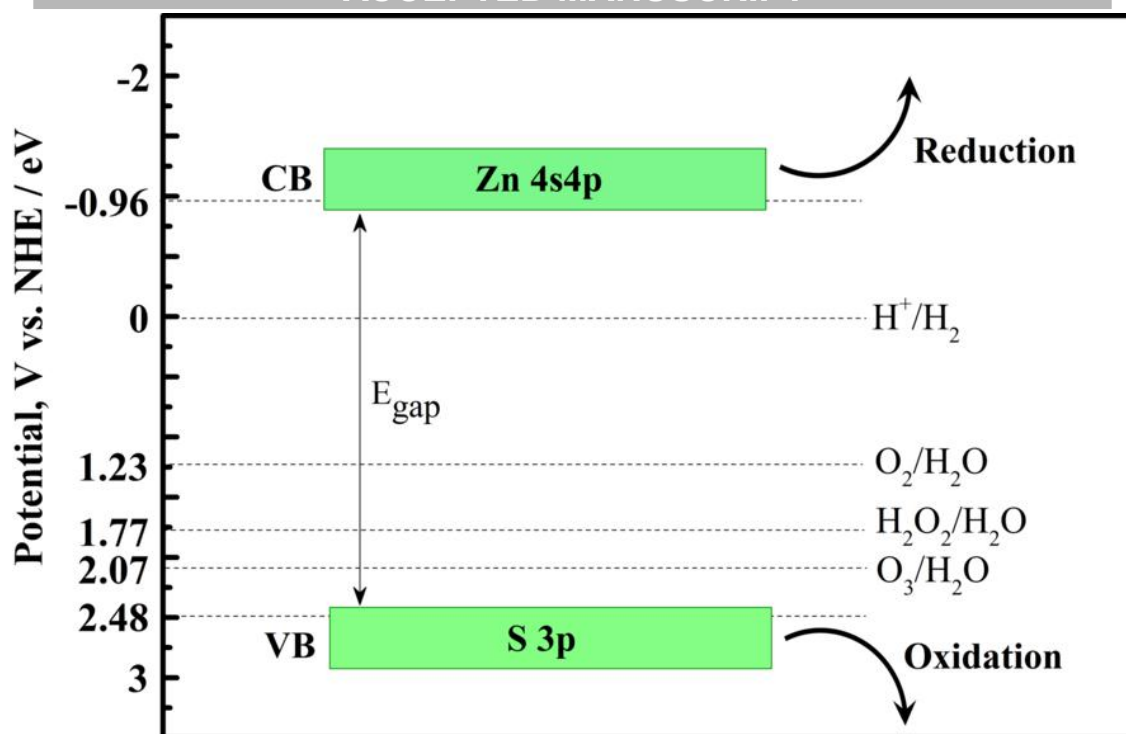


Figure 4.

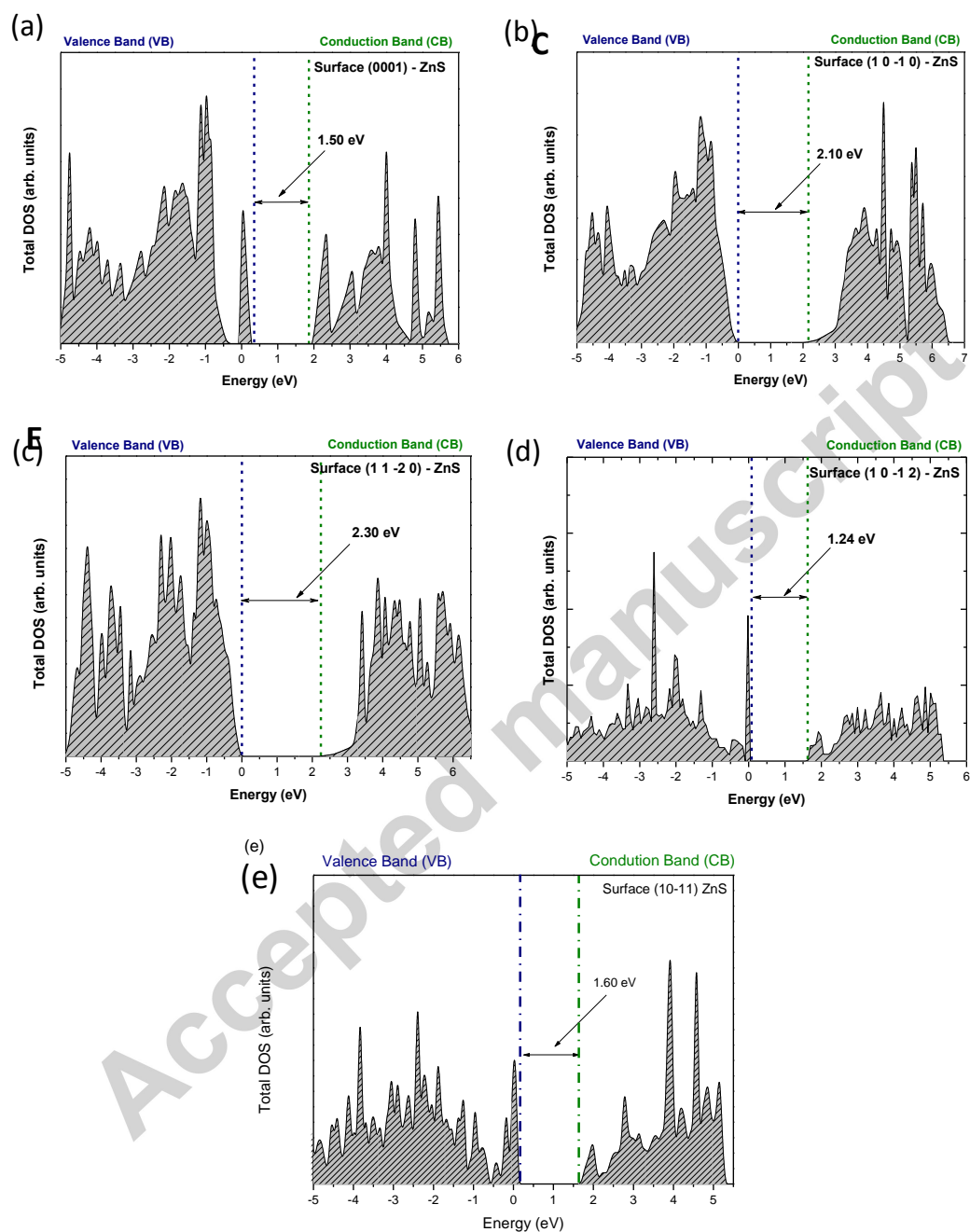


Figure 5.

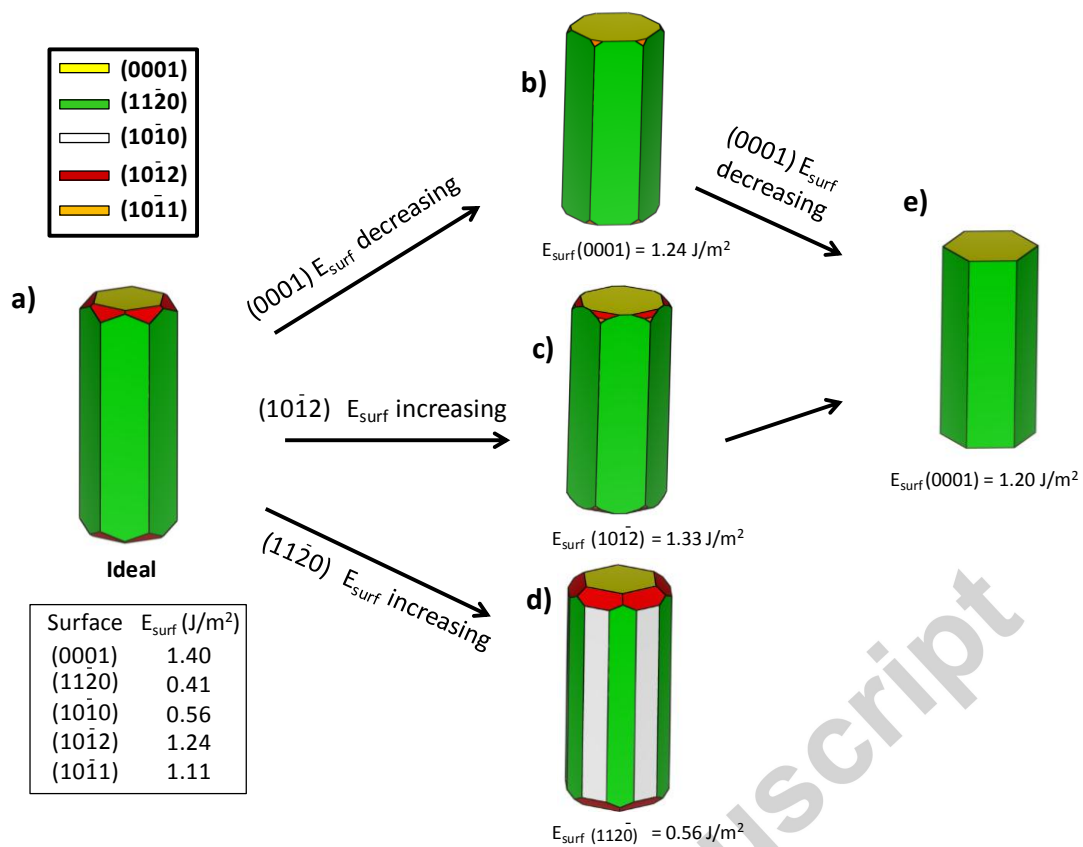


Figure 6.

**Table 1:** Values of  $\rho$  for the ZnS NPs.

Sample	Area 1 (%)	Area 2 (%)	$\rho$
ZN	86.41	53.78	1.61
ZC	85.72	66.30	1.29
ZA	109.57	86.28	1.27

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**Table 2:** Parameter and linear kinetic equation of photocatalytic reaction of the ZnS NPs prepared with different precursors.

Sample	Photodecomposition RhB(%) 210 min	Velocity constant $k \times 10^{-3} (\text{min}^{-1})$	$R^2$	$k/\text{SSA}$ ( $\text{min}^{-1} \cdot \text{m}^{-2} \cdot \text{g}$ )
Without catalysts	12	0.54	0.9548	---
ZA	56	4.08	0.9796	0.0148
ZC	82	7.96	0.9837	0.028
ZN	97	14.05	0.9010	0.060
TiO <sub>2</sub>	44	2.61	0.9560	---



**Table 3:** Number (n) of ZnS units, values of surface area and surface energy ( $E_{\text{surf}}$ ), and the band gap energies ( $E_{\text{gap}}$ ) for (0001), (10 $\bar{1}$ 0), (11 $\bar{2}$ 0) (10 $\bar{1}$ 2) and (10 $\bar{1}$ 1) surfaces of ZnS.

	n	Area (m <sup>2</sup> )	$E_{\text{surf}}$ (J·m <sup>-2</sup> )	$E_{\text{gap}}$ (eV)
(0001)	40	51.3	1.40	1.5
(10 $\bar{1}$ 0)	12	24.3	0.56	2.1
(11 $\bar{2}$ 0)	24	42.0	0.41	2.3
(10 $\bar{1}$ 2)	12	35.3	1.24	1.24
(10 $\bar{1}$ 1)	20	70.6	1.11	1.6

### Highlights

- Effects of different precursors on properties of ZnS nanoparticles have been investigated.
- The optical behavior reveals that the ZnS nanoparticles have potential applications as photocatalytic materials.
- The as-prepared catalyst showed a high activity for the degradation of Rhodamine B